

Lesson	National 5 Chemistry Unit 2.2a Alcohols								Traffic Light					
	Red	Amber	Green											
19	Alcohols are used as fuels as they are highly flammable and burn with very clean flames								☹	☹	☺			
20	Alcohols are often used as solvents <ul style="list-style-type: none"> alcohol is the main constituent of methylated spirits (meths), a useful solvent 								☹	☹	☺			
21	Methanol, ethanol and propanol are miscible with water, thereafter the solubility decreases as size increases <ul style="list-style-type: none"> miscible means that the alcohols mix with water and do not separate. 								☹	☹	☺			
	Alcohol	Methanol	Ethanol	Propanol	Butanol	Pentanol	Hexanol	Heptanol				Octanol		
	No of Carbons	1	2	3	4	5	6	7				8		
	Solubility	Very soluble in water						insoluble						
22	As increase in the size of an alcohol increases the melting & boiling points <ul style="list-style-type: none"> this is caused by the increasing strength of the intermolecular forces. 								☹	☹	☺			
	Alcohol	Methanol	Ethanol	Propanol	Butanol	Pentanol	Hexanol	Heptanol				Octanol		
	No of Carbons	1	2	3	4	5	6	7				8		
	Melting Point	low						high						
	Boiling Point	low						High						
Strength of intermolecular forces	weaker						stronger							
23	An alcohol is a molecule containing a hydroxyl -OH functional group.								☹	☹	☺			
24	Saturated, straight-chain alcohols have the general formula $C_nH_{2n+1}OH$ <ul style="list-style-type: none"> saturated alcohols do not contain $C=C$ double bonds 								☹	☹	☺			
25a 26a	Straight-chain alcohols are named indicating the position of the hydroxyl -OH group can be drawn and named:								☹	☹	☺			
	$\begin{array}{c} H \\ \\ H-C-OH \\ \\ H \end{array}$ methanol	$\begin{array}{c} H & H \\ & \\ H-C-C-OH \\ & \\ H & H \end{array}$ ethanol	$\begin{array}{c} H & H & H \\ & & \\ H-C-C-C-OH \\ & & \\ H & H & H \end{array}$ propan-1-ol	$\begin{array}{c} H & H & H \\ & & \\ H-C-C-C-H \\ & & \\ H & OH & H \end{array}$ propan-2-ol										
	$\begin{array}{c} H & H & H & H \\ & & & \\ H-C-C-C-C-OH \\ & & & \\ H & H & H & H \end{array}$ butan-1-ol	$\begin{array}{c} H & OH & H & H \\ & & & \\ H-C-C-C-C-H \\ & & & \\ H & H & H & H \end{array}$ butan-2-ol	$\begin{array}{c} H & H & H & H & H \\ & & & & \\ H-C-C-C-C-C-OH \\ & & & & \\ H & H & H & H & H \end{array}$ pentan-1-ol	$\begin{array}{c} H & H & H & H & H \\ & & & & \\ H-C-C-C-C-C-H \\ & & & & \\ H & H & H & OH & H \end{array}$ pentan-2-ol										
	$\begin{array}{c} H & H & H & H & H \\ & & & & \\ H-C-C-C-C-C-H \\ & & & & \\ H & H & OH & H & H \end{array}$ pentan-3-ol	$\begin{array}{c} H & H & H & H & H & H \\ & & & & & \\ H-C-C-C-C-C-OH \\ & & & & & \\ H & H & H & H & H & H \end{array}$ hexan-1-ol	$\begin{array}{c} H & H & H & H & H & H \\ & & & & & \\ H-C-C-C-C-C-C-H \\ & & & & & \\ H & H & H & H & OH & H \end{array}$ hexan-2-ol	$\begin{array}{c} H & H & H & H & H & H \\ & & & & & \\ H-C-C-C-C-C-C-H \\ & & & & & \\ H & H & H & OH & H & H \end{array}$ hexan-3-ol										
	Alcohols with branched chains can be drawn:											☹	☹	☺
	$\begin{array}{c} H \\ \\ H-H-C-H \\ & & \\ H-C-C-C-OH \\ & & \\ H & H & H \end{array}$ 2-methylpropan-1-ol	$\begin{array}{c} H \\ \\ H-H-C-H \\ & & \\ H-C-C-C-H \\ & & \\ H & OH & H \end{array}$ 2-methylpropan-2-ol	$\begin{array}{c} H & H & H \\ & & \\ H-C-C-C-H \\ & & \\ H & H & H \end{array}$ 2-methylbutan-1-ol											
	$\begin{array}{c} H \\ \\ H & H-C-H & H & H \\ & & & \\ H-C-C-C-C-H \\ & & & \\ H & OH & H & H \end{array}$ 2-methylbutan-2-ol	$\begin{array}{c} H \\ \\ H & H & H-C-H & H \\ & & & \\ H-C-C-C-C-H \\ & & & \\ H & OH & H & H \end{array}$ 3-methylbutan-2-ol	$\begin{array}{c} H \\ \\ H & H & H-C-H & H \\ & & & \\ H-C-C-C-C-OH \\ & & & \\ H & H & H & H \end{array}$ 3-methylbutan-1-ol											

Nat5 Traffic Lights		Past Paper Question Bank Unit 2.2a Alcohols										JABchem				
Outcome	Original Specimen Paper	New Specimen Paper	Nat5 2014	Nat5 2015	Nat5 2016	Nat5 2017	Nat5 2018	Nat5 2019	Nat5 2020	Nat5 2021						
19																
20																
21																
22			mc13				mc14									
23	L8a	L8a	L8b(i)	L3a(i)	L12a	L12a		L2a								
24																
25a 26a						L14a(i)	mc13									
25b 26b			L8b(ii)	mc13												

Nat5	Answer	% Correct	Reasoning																																																										
2014 MC 13	C	88	<table border="1"> <thead> <tr> <th>Alkanol</th> <th>Propan-1-ol</th> <th>Propan-2-ol</th> <th>Butan-1-ol</th> <th>Butan-2-ol</th> </tr> </thead> <tbody> <tr> <td>Boiling Point</td> <td>97°C</td> <td>82°C</td> <td>118°C</td> <td>100°C</td> </tr> </tbody> </table>					Alkanol	Propan-1-ol	Propan-2-ol	Butan-1-ol	Butan-2-ol	Boiling Point	97°C	82°C	118°C	100°C																																												
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2015 MC 13	A	81	<input checked="" type="checkbox"/> A correct structure is drawn. <input checked="" type="checkbox"/> B Shortened structural formula would be: $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}(\text{OH})\text{C}(\text{CH}_3)_3$ <input checked="" type="checkbox"/> C Shortened structural formula would be: $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}(\text{OH})\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_3$ <input checked="" type="checkbox"/> D Shortened structural formula would be: $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}(\text{OH})\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$																																																										
2018 MC 13	A	-	<input checked="" type="checkbox"/> A pentan-2-ol: five carbons linked by single bonds with -OH group in Carbon No 2 <input checked="" type="checkbox"/> B Pentan-4-ol: The -OH group has not given the lowest numbering system <input checked="" type="checkbox"/> C 1-methylbutan-3-ol: Longest continuous chain of carbons with -OH group is 5 <input checked="" type="checkbox"/> D 4-methylbutan-1-ol Longest continuous chain of carbons with -OH group is 5																																																										
2018 MC 14	B	-	<table border="1"> <thead> <tr> <th>Alcohol</th> <th>Methanol</th> <th>Ethanol</th> <th>Propanol</th> <th>Butanol</th> <th>Pentanol</th> <th>Hexanol</th> <th>Heptanol</th> <th>Octanol</th> </tr> </thead> <tbody> <tr> <td>Formula</td> <td>CH_3OH</td> <td>$\text{C}_2\text{H}_5\text{OH}$</td> <td>$\text{C}_3\text{H}_7\text{OH}$</td> <td>$\text{C}_4\text{H}_9\text{OH}$</td> <td>$\text{C}_5\text{H}_{11}\text{OH}$</td> <td>$\text{C}_6\text{H}_{13}\text{OH}$</td> <td>$\text{C}_7\text{H}_{15}\text{OH}$</td> <td>$\text{C}_8\text{H}_{17}\text{OH}$</td> </tr> <tr> <td>Melting Point</td> <td>low</td> <td colspan="6" style="text-align: center;">—————→</td> <td>high</td> </tr> <tr> <td>Boiling Point</td> <td>low</td> <td colspan="6" style="text-align: center;">—————→</td> <td>high</td> </tr> <tr> <td>Solubility</td> <td>high</td> <td colspan="6" style="text-align: center;">←—————</td> <td>low</td> </tr> <tr> <td>Strength of Intermolecular Bonds</td> <td>low</td> <td colspan="6" style="text-align: center;">—————→</td> <td>high</td> </tr> </tbody> </table>					Alcohol	Methanol	Ethanol	Propanol	Butanol	Pentanol	Hexanol	Heptanol	Octanol	Formula	CH_3OH	$\text{C}_2\text{H}_5\text{OH}$	$\text{C}_3\text{H}_7\text{OH}$	$\text{C}_4\text{H}_9\text{OH}$	$\text{C}_5\text{H}_{11}\text{OH}$	$\text{C}_6\text{H}_{13}\text{OH}$	$\text{C}_7\text{H}_{15}\text{OH}$	$\text{C}_8\text{H}_{17}\text{OH}$	Melting Point	low	—————→						high	Boiling Point	low	—————→						high	Solubility	high	←—————						low	Strength of Intermolecular Bonds	low	—————→						high
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Nat5	Answer	Reasoning									
2014 8b(i)	hydroxyl	The -OH group of the hydroxyl group.									
2014 8b(ii)	Any structure from:	<p>Correct C₅H₁₁OH diagram of:</p> <table border="1"> <tr> <td>pentan-1-ol</td> <td>pentan-2-ol</td> <td>pentan-3-ol</td> </tr> <tr> <td colspan="2">2-methylbutan-1-ol</td> <td>2,2-dimethylpropan-1-ol</td> </tr> <tr> <td colspan="2">2-methylbutan-2-ol</td> <td>3-methylbutan-2-ol</td> </tr> </table> <p>NB: diagram must be different from 3-methylbutan-1-ol in question and not a redrawing of same 3-methylbutan-1-ol structure.</p>	pentan-1-ol	pentan-2-ol	pentan-3-ol	2-methylbutan-1-ol		2,2-dimethylpropan-1-ol	2-methylbutan-2-ol		3-methylbutan-2-ol
pentan-1-ol	pentan-2-ol	pentan-3-ol									
2-methylbutan-1-ol		2,2-dimethylpropan-1-ol									
2-methylbutan-2-ol		3-methylbutan-2-ol									
2015 3a(i)	hydroxyl	Hydroxyl groups are the functional group found in all alcohols and have the structure -OH or -O-H									
2016 12a	Hydroxyl group	<table border="1"> <tr> <td style="text-align: center;">-O-H</td> <td style="text-align: center;"> $\begin{array}{c} \text{O} \\ \\ -\text{C}-\text{OH} \end{array}$ </td> </tr> <tr> <td style="text-align: center;">hydroxyl group</td> <td style="text-align: center;">carboxyl group</td> </tr> </table>	-O-H	$\begin{array}{c} \text{O} \\ \\ -\text{C}-\text{OH} \end{array}$	hydroxyl group	carboxyl group					
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hydroxyl group	carboxyl group										
2017 12a	C=C double bond or -OH group	C=C double bonds are the functional groups found in alkenes Hydroxyl -OH groups are the functional group found in alcohols									
2017 14a(i)	$\begin{array}{ccccccc} \text{H} & \text{H} & \text{H} & \text{H} & \text{H} & \text{H} & \\ & & & & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & & & & \\ \text{H} & \text{H} & \text{H} & \text{H} & \text{H} & \text{H} & \end{array}$	<table border="1"> <tr> <td style="text-align: center;"> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{OH} \\ \\ \text{H} \end{array}$ Methanol </td> <td style="text-align: center;"> $\begin{array}{c} \text{H} & \text{H} \\ & \\ \text{H}-\text{C}-\text{C}-\text{OH} \\ & \\ \text{H} & \text{H} \end{array}$ Ethanol </td> <td style="text-align: center;"> $\begin{array}{c} \text{H} & \text{H} & \text{H} \\ & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & \\ \text{H} & \text{H} & \text{H} \end{array}$ Propan-1-ol </td> </tr> <tr> <td style="text-align: center;"> $\begin{array}{c} \text{H} & \text{H} & \text{H} & \text{H} \\ & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & \\ \text{H} & \text{H} & \text{H} & \text{H} \end{array}$ Butan-1-ol </td> <td style="text-align: center;"> $\begin{array}{c} \text{H} & \text{H} & \text{H} & \text{H} & \text{H} \\ & & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & & \\ \text{H} & \text{H} & \text{H} & \text{H} & \text{H} \end{array}$ Pentan-1-ol </td> <td style="text-align: center;"> $\begin{array}{c} \text{H} & \text{H} & \text{H} & \text{H} & \text{H} & \text{H} \\ & & & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & & & \\ \text{H} & \text{H} & \text{H} & \text{H} & \text{H} & \text{H} \end{array}$ Hexan-1-ol </td> </tr> </table>	$\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{OH} \\ \\ \text{H} \end{array}$ Methanol	$\begin{array}{c} \text{H} & \text{H} \\ & \\ \text{H}-\text{C}-\text{C}-\text{OH} \\ & \\ \text{H} & \text{H} \end{array}$ Ethanol	$\begin{array}{c} \text{H} & \text{H} & \text{H} \\ & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & \\ \text{H} & \text{H} & \text{H} \end{array}$ Propan-1-ol	$\begin{array}{c} \text{H} & \text{H} & \text{H} & \text{H} \\ & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & \\ \text{H} & \text{H} & \text{H} & \text{H} \end{array}$ Butan-1-ol	$\begin{array}{c} \text{H} & \text{H} & \text{H} & \text{H} & \text{H} \\ & & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & & \\ \text{H} & \text{H} & \text{H} & \text{H} & \text{H} \end{array}$ Pentan-1-ol	$\begin{array}{c} \text{H} & \text{H} & \text{H} & \text{H} & \text{H} & \text{H} \\ & & & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & & & \\ \text{H} & \text{H} & \text{H} & \text{H} & \text{H} & \text{H} \end{array}$ Hexan-1-ol			
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2019 2a	Hydroxyl	<table border="1"> <tr> <td style="text-align: center;">-O-H</td> <td style="text-align: center;"> $\begin{array}{c} \text{O} \\ \\ -\text{C}-\text{OH} \end{array}$ </td> </tr> <tr> <td style="text-align: center;">hydroxyl group</td> <td style="text-align: center;">carboxyl group</td> </tr> </table>	-O-H	$\begin{array}{c} \text{O} \\ \\ -\text{C}-\text{OH} \end{array}$	hydroxyl group	carboxyl group					
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Nat5 Traffic Lights		Past Paper Question Bank										JABchem				
		Unit 2.2a Alcohols														
Outcome	Int2 2000	Int2 2001	Int2 2002	Int2 2003	Int2 2004	Int2 2005	Int2 2006	Int2 2007	Int2 2008	Int2 2009	Int2 2010	Int2 2011	Int2 2012	Int2 2013	Int2 2014	Int2 2015
19																
20																
21																
22			mc11													
23			L5b(ii)			mc11			L6a		L9c			L13a(ii)		
24																
25a 26a						L6a(i)	L7a					L9c	L7c			L6a
25b 26b																

Int2	Answer	% Correct	Reasoning															
2002 MC 11	C	78	<table border="1"> <thead> <tr> <th>Substance</th> <th>A</th> <th>B</th> <th>C</th> <th>D</th> </tr> </thead> <tbody> <tr> <td>Name</td> <td>Propan-1-ol</td> <td>Propan-2-ol</td> <td>Butan-1-ol</td> <td>Butan-2-ol</td> </tr> <tr> <td>Boiling Point (°C)</td> <td>97</td> <td>82</td> <td>117</td> <td>100</td> </tr> </tbody> </table>	Substance	A	B	C	D	Name	Propan-1-ol	Propan-2-ol	Butan-1-ol	Butan-2-ol	Boiling Point (°C)	97	82	117	100
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Name	Propan-1-ol	Propan-2-ol	Butan-1-ol	Butan-2-ol														
Boiling Point (°C)	97	82	117	100														
2005 MC 11	B	66	<input checked="" type="checkbox"/> A Molecule shown has C=C double bond so it is unsaturated <input checked="" type="checkbox"/> B Molecule is unsaturated (C=C double bond) and an alcohol (contains -OH group) <input checked="" type="checkbox"/> C Molecule shown has C=C double bond so it is unsaturated <input checked="" type="checkbox"/> D Molecule shown has hydroxyl group so it is an alcohol not a carboxylic acid															

Int2	Answer	Reasoning						
2002 5b(i)	$-O-H$	Alcohols contain the hydroxyl $-OH$ functional group						
2005 6a(i)	$ \begin{array}{c} H \quad OH \quad H \\ \quad \quad \\ H-C-C-C-H \\ \quad \quad \\ H \quad H \quad H \end{array} $	<p style="text-align: center;">propan-2-ol</p> <p style="text-align: center;"> </p>						
2006 7a	$ \begin{array}{c} H \quad H \\ \quad \\ H-C-C-OH \\ \quad \\ H \quad H \end{array} $	<p><i>Ethanol</i> is a 2 carbon structure.</p> <p><i>Ethanol</i> is an alcohol with a hydroxyl $-OH$ functional group</p>						
2008 6a	hydroxyl	The hydroxyl group is the $-OH$ group						
2010 9c	$-OH$ group circled	The hydroxyl functional group has the formula $-O-H$						
2011 9c	$ \begin{array}{c} H \quad H \\ \quad \\ H-C-C-OH \\ \quad \\ H \quad H \end{array} $	<p><u>Ethanol</u> is a 2 carbon structure.</p> <p><u>Ethanol</u> is an alcohol with a hydroxyl $-OH$ functional group</p>						
2012 7c	One from:	<table border="1" style="width: 100%; text-align: center;"> <thead> <tr> <th style="width: 33%;">Butan-1-ol C_4H_9OH</th> <th style="width: 33%;">2-methylpropan-1-ol C_4H_9OH</th> <th style="width: 33%;">2-methylpropan-2-ol C_4H_9OH</th> </tr> </thead> <tbody> <tr> <td> $\begin{array}{c} H \quad H \quad H \quad H \\ \quad \quad \quad \\ H-C-C-C-C-OH \\ \quad \quad \quad \\ H \quad H \quad H \quad H \end{array}$ </td> <td> $\begin{array}{c} H \\ \\ H-H-C-H \\ \quad \quad \\ H-C-C-C-OH \\ \quad \quad \\ H \quad H \quad H \end{array}$ </td> <td> $\begin{array}{c} H \\ \\ H-H-C-H \\ \quad \quad \\ H-C-C-C-H \\ \quad \quad \\ H \quad OH \quad H \end{array}$ </td> </tr> </tbody> </table>	Butan-1-ol C_4H_9OH	2-methylpropan-1-ol C_4H_9OH	2-methylpropan-2-ol C_4H_9OH	$ \begin{array}{c} H \quad H \quad H \quad H \\ \quad \quad \quad \\ H-C-C-C-C-OH \\ \quad \quad \quad \\ H \quad H \quad H \quad H \end{array} $	$ \begin{array}{c} H \\ \\ H-H-C-H \\ \quad \quad \\ H-C-C-C-OH \\ \quad \quad \\ H \quad H \quad H \end{array} $	$ \begin{array}{c} H \\ \\ H-H-C-H \\ \quad \quad \\ H-C-C-C-H \\ \quad \quad \\ H \quad OH \quad H \end{array} $
Butan-1-ol C_4H_9OH	2-methylpropan-1-ol C_4H_9OH	2-methylpropan-2-ol C_4H_9OH						
$ \begin{array}{c} H \quad H \quad H \quad H \\ \quad \quad \quad \\ H-C-C-C-C-OH \\ \quad \quad \quad \\ H \quad H \quad H \quad H \end{array} $	$ \begin{array}{c} H \\ \\ H-H-C-H \\ \quad \quad \\ H-C-C-C-OH \\ \quad \quad \\ H \quad H \quad H \end{array} $	$ \begin{array}{c} H \\ \\ H-H-C-H \\ \quad \quad \\ H-C-C-C-H \\ \quad \quad \\ H \quad OH \quad H \end{array} $						
2013 13a(ii)	$-OH$ group (bottom on the right)	<p>Hydroxyl groups have the formula $-OH$</p> <p>Carboxyl $-COOH$ groups contain an $-OH$ group within the structure but the proximity of the $C=O$ group to the $-OH$ group changes the properties of the $-OH$ group to the properties of the carboxyl $-COOH$ group.</p>						
2015 6a	Propan-2-ol	<p style="text-align: center;">Propan -2- ol</p> <p style="text-align: center;"> </p>						

Nat5
Traffic Lights

Past Paper Question Bank

Unit 2.2a Alcohols

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Outcome	2000 Credit	2001 Credit	2002 Credit	2003 Credit	2004 Credit	2005 Credit	2006 Credit	2007 Credit	2008 Credit	2009 Credit	2010 Credit	2011 Credit	2012 Credit	2013 Credit		
19																
20																
21																
22																
23																
24															15a	
25a 26a											20a(i)					
25b 26b																

SG Credit	Answer	Reasoning																		
2010C 20a(i)	Propanol	$\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{OH} \\ \\ \text{H} \end{array}$ methanol	$\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{OH} \\ \quad \\ \text{H} \quad \text{H} \end{array}$ ethanol	$\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{OH} \\ \quad \quad \\ \text{H} \quad \text{H} \quad \text{H} \end{array}$ propanol	$\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \quad \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ \quad \quad \quad \\ \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \end{array}$ butanol															
2013C 15a	$C_nH_{2n+1}OH$ or $C_nH_{2n+2}O$	<table border="1" style="width: 100%;"> <thead> <tr> <th colspan="2">Ethanol</th> <th colspan="2">Propan-2-ol</th> </tr> </thead> <tbody> <tr> <td>If formula written as: C_2H_5OH</td> <td>Molecular formula: C_2H_6O</td> <td>If formula written as: C_3H_7OH</td> <td>Molecular formula: C_3H_8O</td> </tr> <tr> <td>If $n=2$ then $2n+1=5$</td> <td>If $n=2$ then $2n+2=6$</td> <td>If $n=3$ then $2n+1=7$</td> <td>If $n=3$ then $2n+2=8$</td> </tr> <tr> <td>$\therefore C_nH_{2n+1}OH$</td> <td>$\therefore C_nH_{2n+2}O$</td> <td>$\therefore C_nH_{2n+1}OH$</td> <td>$\therefore C_nH_{2n+2}O$</td> </tr> </tbody> </table>		Ethanol		Propan-2-ol		If formula written as: C_2H_5OH	Molecular formula: C_2H_6O	If formula written as: C_3H_7OH	Molecular formula: C_3H_8O	If $n=2$ then $2n+1=5$	If $n=2$ then $2n+2=6$	If $n=3$ then $2n+1=7$	If $n=3$ then $2n+2=8$	$\therefore C_nH_{2n+1}OH$	$\therefore C_nH_{2n+2}O$	$\therefore C_nH_{2n+1}OH$	$\therefore C_nH_{2n+2}O$	
Ethanol		Propan-2-ol																		
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$\therefore C_nH_{2n+1}OH$	$\therefore C_nH_{2n+2}O$	$\therefore C_nH_{2n+1}OH$	$\therefore C_nH_{2n+2}O$																	

Past Paper Question Bank

Unit 2.2a Alcohols

Outcome	2000 <i>General</i>	2001 <i>General</i>	2002 <i>General</i>	2003 <i>General</i>	2004 <i>General</i>	2005 <i>General</i>	2006 <i>General</i>	2007 <i>General</i>	2008 <i>General</i>	2009 <i>General</i>	2010 <i>General</i>	2011 <i>General</i>	2012 <i>General</i>	2013 <i>General</i>		
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